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Explain why ethanol ( $\text{CH}_3\text{CH}_2\text{OH}$ )

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is more soluble in water than is ethane ( $\text{CH}_3\text{CH}_3$ ). Answer Ethanol is polar; ethane is not. The ethanol  $-\text{OH}$  group can hydrogen-bond with water. 2.

Calculation of pH from Hydrogen Ion Concentration  
What is the pH of a solution that has an  $\text{H}^+$  concentration of  
(a)  $1.75 \times 10^{-5} \text{ mol/L}$ ; (b)  $6.50 \times 10^{-10} \text{ mol/L}$ ; (c)  $1.0 \times 10^{-4} \text{ mol/L}$ ;

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