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102 at Maseno University. 1. Base your
answer to the following question on
Given the nuclear equation: $X + {}^6_3\text{Li} \rightarrow {}^7_4\text{Be} + {}^4_2\text{He}$ The

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Example 2: Uranium-239 decays by emitting a beta and a γ , Hour:
Page 3 ${}^{239}_{92}\text{U} \rightarrow e^{-} + \gamma + {}^{239}_{93}\text{Np}$ NOTE: notice that $92 - [-1] = 93$; there is always an increase in the atomic number with beta emission. In the above examples you should notice that the sum of the masses on the left of the arrow equals the sum of

I,

Name Unit 11: Nuclear Chemistry
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1. Base your answer to the following question on Given the nuclear equation:
 ${}^{1}_{1}\text{H} + X \rightarrow {}^{6}_{3}\text{Li} + {}^{4}_{2}\text{He}$ A) ${}^{94}_{11}\text{Li}$ B) ${}^{94}_{10}\text{Be}$
C) ${}^{105}_{11}\text{Be}$ D) ${}^{106}_{12}\text{C}$ The particle represented by X is A) ${}^{53}_{26}\text{Fe}$ B) ${}^{137}_{55}\text{Cs}$
C) ${}^{198}_{79}\text{Au}$ D) ${}^{220}_{87}\text{Fr}$ 2. Which isotope will

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Answers

spontaneously decay and emit particles with a charge of +2? A)alpha B ...

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14. 23. 4. Key Ideas. Stability of isotopes is based in the ratio of neutrons and protons in its nucleus. Although most nuclei are stable, some are unstable and spontaneously decay, emitting radiation.

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Answers

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1. Base your answer to the following question on Given the nuclear equation:
 $11\text{H} + \text{X} \rightarrow 63\text{Li} + 42\text{He}$ A) 94Li B) 94Be
C) 105Be D) 106C The particle represented by X is

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1. Base your answer to the following question on Given the nuclear equation:
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C) 105Be D) 106C The particle represented by X is A) 53Fe B) 137Cs
C) 198Au D) 220Fr 2. Which isotope will spontaneously decay and emit particles

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