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Chapter 14 The
Behavior of Gases 147

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SECTION 14.1

PROPERTIES OF

GASES(pages 413-417)

This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature.

Compressibility (pages 413-414) 1. Look at Figure 14.1 on page 413.

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SECTION 14.1
PROPERTIES OF
GASES(pages
413-417)

The characteristic or properties of gases to fill the available volume within a container is the result of the freedom that gas particles have to move everywhere in the accessible space. This autonomy of movement of gaseous molecules is because of the very weak

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binding forces amidst
molecules.

What are the Properties of Gases? - Physical Properties Of ...

Properties of Matter:
Gases. By Mary Bagley
- LiveScience

Contributor 08 January
2016. ... The units
relate to one another
this way: 1 atm = 14.7
psi = 760 mmHg =
101.3 kPa (1,000
pascals ...

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Properties of Matter: Gases | Live Science

14.1 Properties of
Gases In organized
soccer, there are rules
about equipment. For
international
competitions, the ball's
mass must be not
more than 450 grams
and not less than 410
grams. The pressure of
the air inside the ball
must be no lower than
0.6 atmospheres and

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no higher than

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The noble gases and mercury occur as monatomic species, whereas all other

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gases and bromine are diatomic molecules. All of the gaseous elements (other than the monatomic noble gases) are molecules. Within the same group (1, 15, 16 and 17), the lightest elements are gases.

6.1: Properties of Gases: Gas Pressure - Chemistry LibreTexts

Section 14.1 The
Properties of Gases.

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OBJECTIVES: Explain why gases are easier to compress than solids or liquids are. *

Section 14.1 The Properties of Gases.

OBJECTIVES: Describe the three factors that affect gas pressure. *

Compressibility. Gases can expand to fill its container, unlike solids or liquids;

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Behavior of Gases -
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Chapter 14: The
Properties of Gases
Gas Properties
Questions 1. Why do
you suppose the
pressure of
atmospheric gases
increases with
proximity to earth?
Why aren't the gases
evenly distributed
throughout the entire
reaches of the
atmosphere? Study the
PhET simulation for
further illustration
here:

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Chemistry Student Edition - Basic Answer Key Chapter 14 ...

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Chapter 14 Properties
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SECTION 14.1
PROPERTIES OF
GASES(pages 413-417)

14.1 Properties of
Gases In organized
soccer, there are rules
about equipment. For
international
competitions, the ball's

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mass must be not more than 450 grams and not less than 410 grams.

Chapter 14 Properties Of Gases Answers

Q2.14. What volume of air at 1.0 atm and 22°C is needed to fill a 0.98 L bicycle tire to a pressure of 5.0 atm at the same temperature? (Note that the 5.0 atm is the gauge pressure, which is the difference

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between the pressure in the tire and the atmospheric pressure.

2.E: Properties of Gases (Exercises) - Chemistry LibreTexts

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CDS 14 Physical Properties of Gases I.
... So our goal here is going to be to relate the properties of gases to each other and then in the subsequent

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concept development study we're going to do the work of relating these properties to the properties of the individual gas molecules.

CDS 14 Physical Properties of Gases I - Ideal Gas Law and

...

14.1 Properties of
Gases 14 Chapter 14
The Properties of
Gases Crossword Down
: 1) The pressure of a

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gas is directly proportional to the Kelvin temperature if the volume remains constant. 2) The relationship $PV=nRT$, which describes the behavior of an ideal gas.

Chapter 14 Properties Of Gases Answers

Chapter 14 Properties of Gases. The Properties of Gases. Gas can expand to fill

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its container Gases are easily compressed, or squeezed into a smaller volume. Gases occupy far more space than a liquid or a solid Slideshow 301683 by fawzia

PPT - Chapter 14 Properties of Gases PowerPoint ...

In lecture, we're going to continue our study of the physical properties of gases, building on the

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material and concept development study number 14. This will also build upon the material from lecture one, in which we studied the relationship between pressure of the gas and the volume of gas for fixed samples of gas at fixed temperatures.

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